

GEF4400 “The Earth System”

Prof. Dr. Jon Egill Kristjansson,

Prof. Dr. Kirstin Krüger (UiO)

- **Lecture/ interactive seminar/ field excursion**

Teaching language: English

Time and location: Monday 12:15-14:00

Thursday 14:15-16:00, CIENS Glasshallen 2.

- **Study program**

Master of meteorology and oceanography

PhD course for meteorology and oceanography students

- **Credits and conditions:**

The successful completion of the course includes an **oral presentation (weight 50%)**, a **successful completion of the Andøya field excursion (mandatory)**, a **field report**, as well as a final **oral examination (50%)**. Student presentations will be part of the course.



GEF4400/9400 **changed** time schedule

Changed GEF4400/9400 time schedule during November 2015:

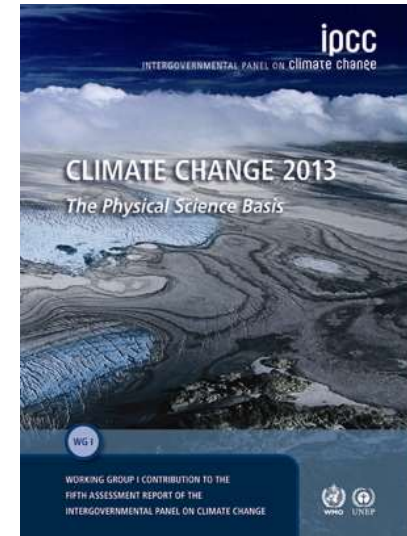
Mo. 02.11.15: **10:00-12:30**, Wed 04.11.15 10:15-12:00

Mo. 09.11.15: **10:00-12:30**, Wed 11.11.15 10:15-12:00

Mo. 16.11.15: **10:00-12:30**, Wed 18.11.15 10:15-12:00

Mo. 23.11.15: **10:00-12:30**, Wed 25.11.15 10:15-12:00

IPCC Chapter 6: Carbon and other biogeochemical cycles



- Background
- Introduction: Global Carbon Cycle (*Section 6.1*)
- Variations in Carbon cycle before the fossil fuel era (*Section 6.2*)
- Evolution of biogeochemical cycles since industrial era (*Section 6.3*)
- Projections of future carbon cycles (*Section 6.4*)
- Executive Summary (Ch. 6)

Ciais, P., et al., 2013: Carbon and Other Biogeochemical Cycles. In: Climate Change 2013: The Physical Science Basis. Contribution of Working Group I to the Fifth Assessment Report of the Intergovernmental Panel on Climate Change. Cambridge University Press.

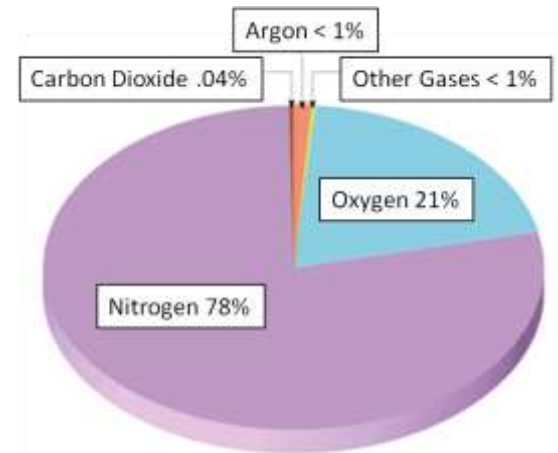


Carbon dioxide (CO₂) in the atmosphere

Content of air

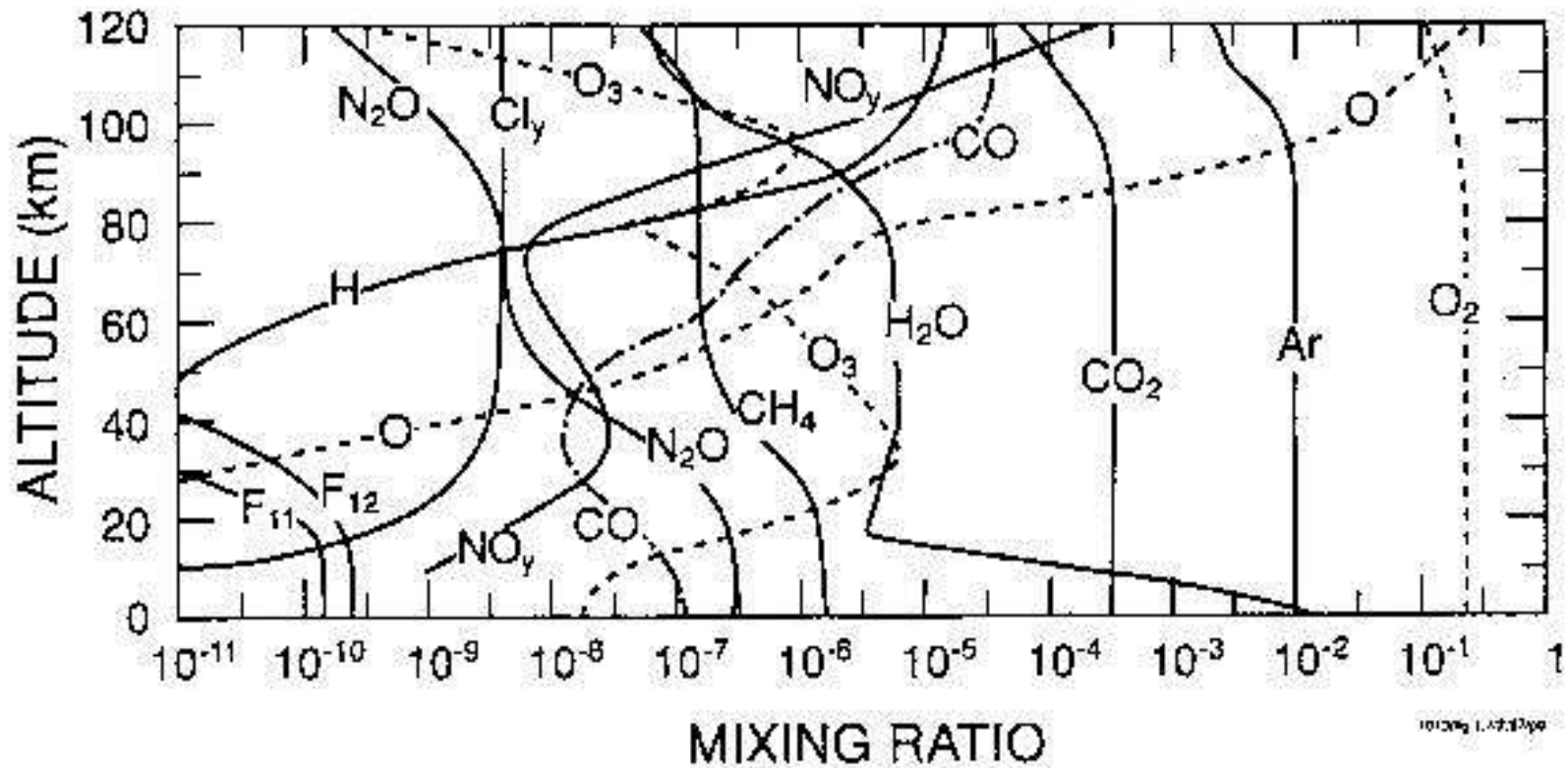
Nitrogen N ₂	78
Oxygen O ₂	21
Inert gases (Ar)	0.9
	<hr/>
	99.9

Percent of volume (%)



Carbon dioxide CO ₂	0.04, varying
Ozone O ₃	0.00005, varying
Water vapour H ₂ O	highly varying
+ other trace gases	

Background



Typical vertical distribution of chemical species within the air [Brasseur, 1999].

Mixing ratios of trace gases

1 ppm (1 part per million) 1 particle CO₂ per 10⁶ particles air

1 ppb (1 part per billion): 1 particle CO₂ per 10⁹ particles air

1 ppt (1 part per trillion): 1 particle CO₂ per 10¹² particles air

“v”: per volume

“m”: per mass

Mixing ratio is a relative unit → taking the air density into account

Absolute unit: concentration of a trace gas (e.g. given in mPa, nbar,
Dobson units for ozone)

Carbon dioxide

Molecular formula

CO₂

Molar mass

44.010 g/mol

Appearance

colorless, odorless gas

(gas at 1 atm and 0 °C; 1 atm =1013.25 hPa)

(solid at 1 atm and -78.5 °C)

(liquid at 56 atm and 20 °C)

Dipole moment

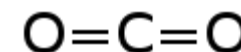
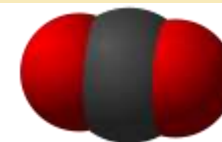
zero

Molecular shape

linear

Spectral data

UV, IR

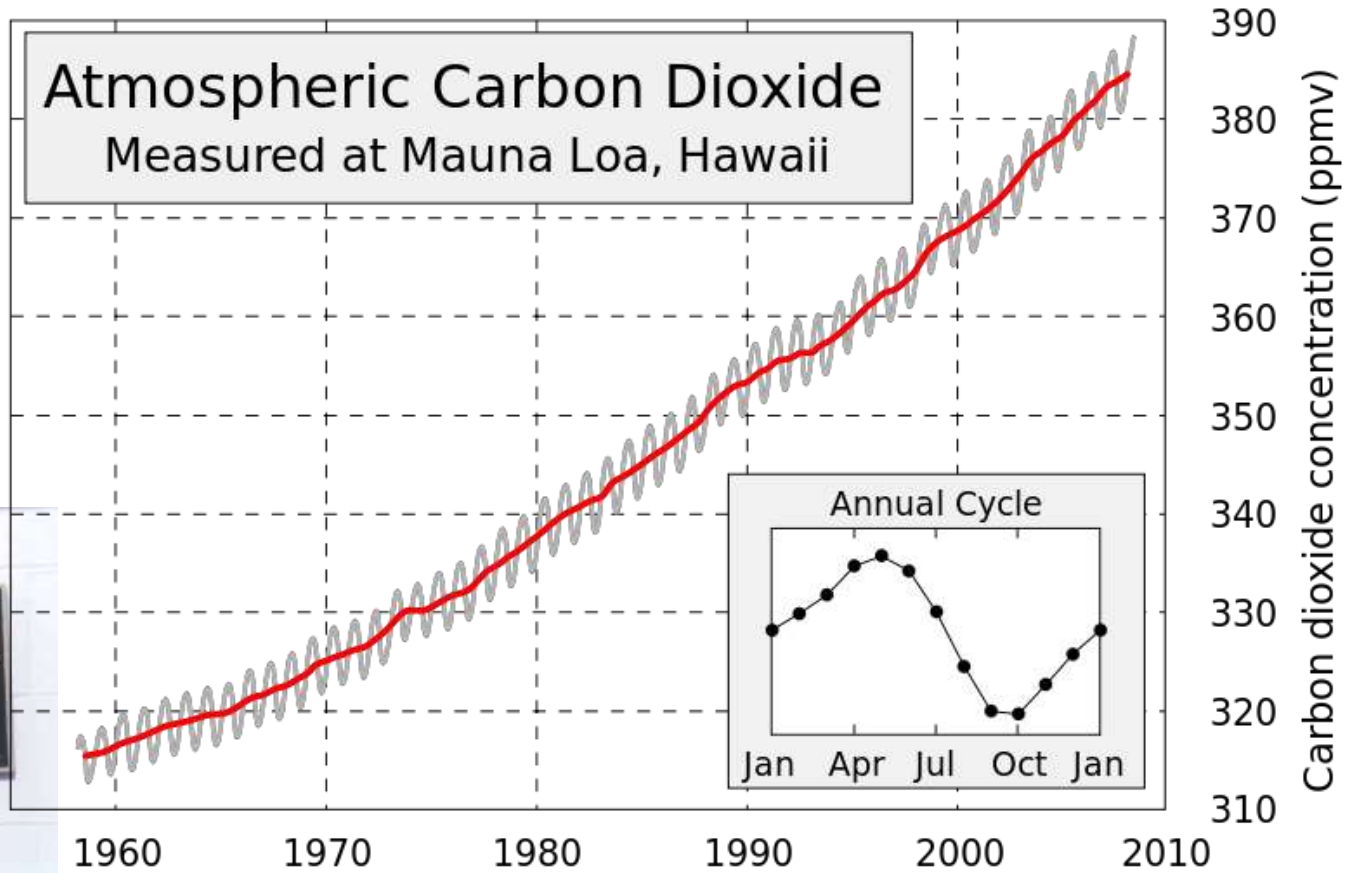


CO₂: **trace gas, 0.038%** concentration of the atmosphere,
sources and sinks are at the surface, uniform distribution up to 90 km.

Sources: combustion of fossil fuels, burning of vegetable matter, chemical processes, respiration, volcanoes , geothermal processes, dissolution of carbonates in crustal rocks.

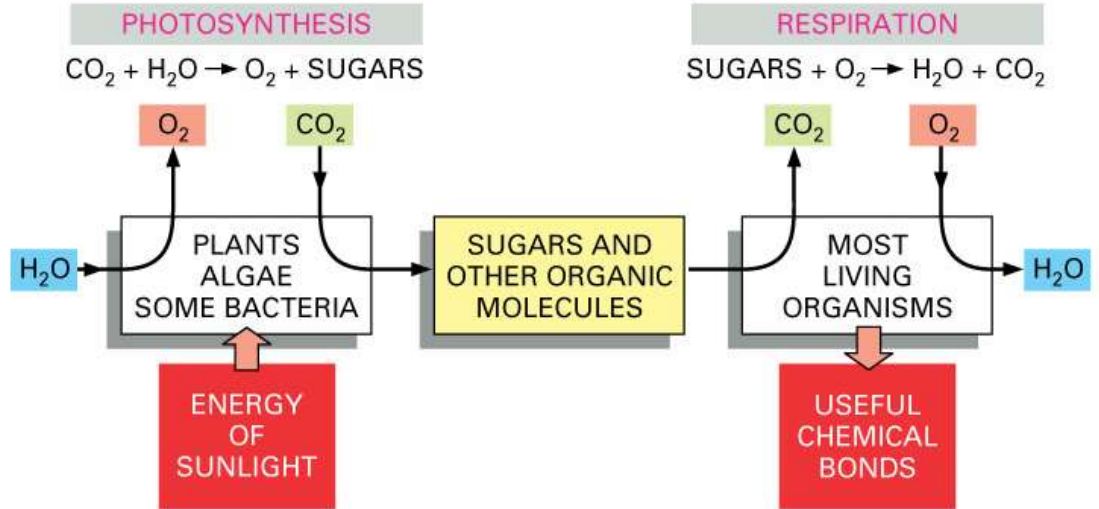
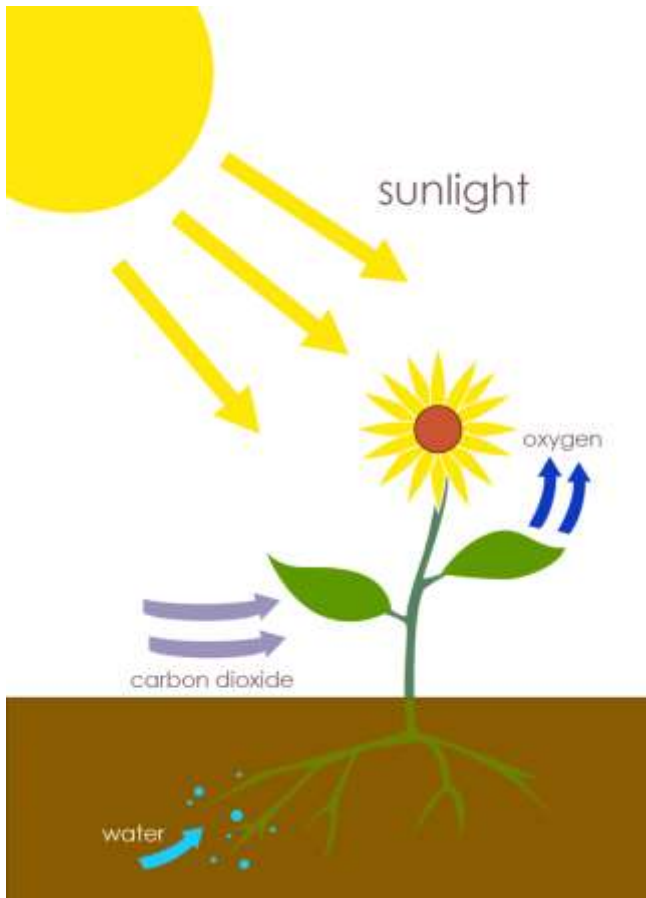
Sinks: ocean, sediments, biosphere(photosynthesis)

Mauna loa curve (Keeling curve)

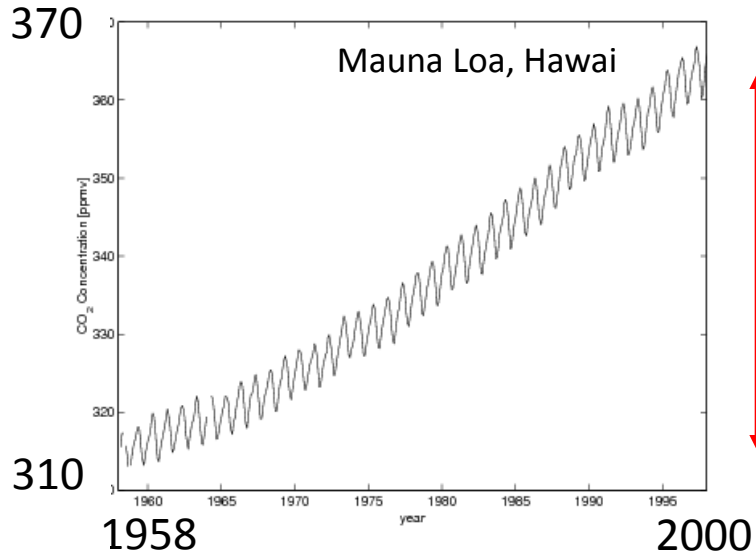


Charles D. Keeling
1928-2005

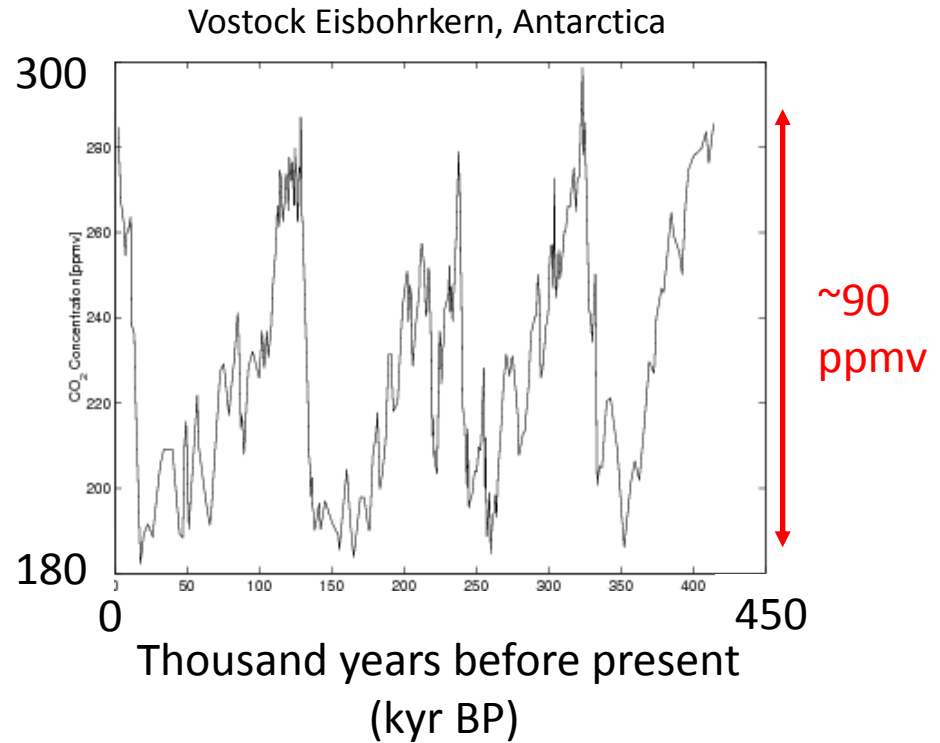
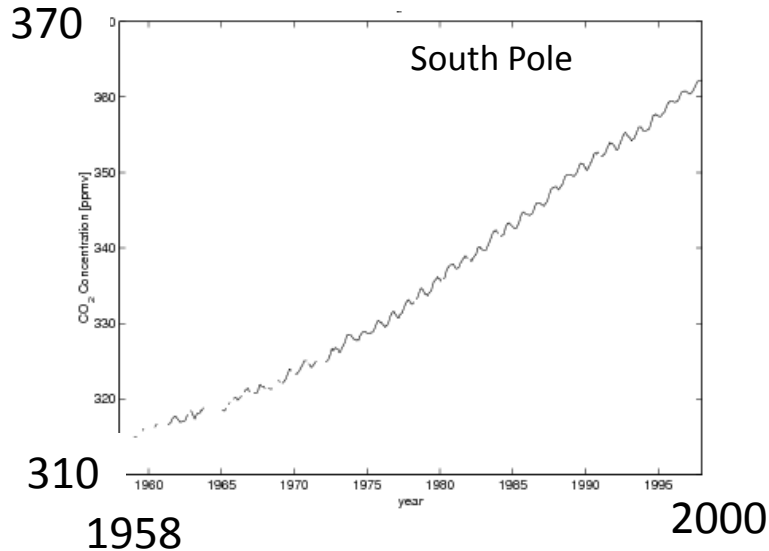
Photosynthesis + Respiration



Past and present CO₂ mixing ratio (ppmv)



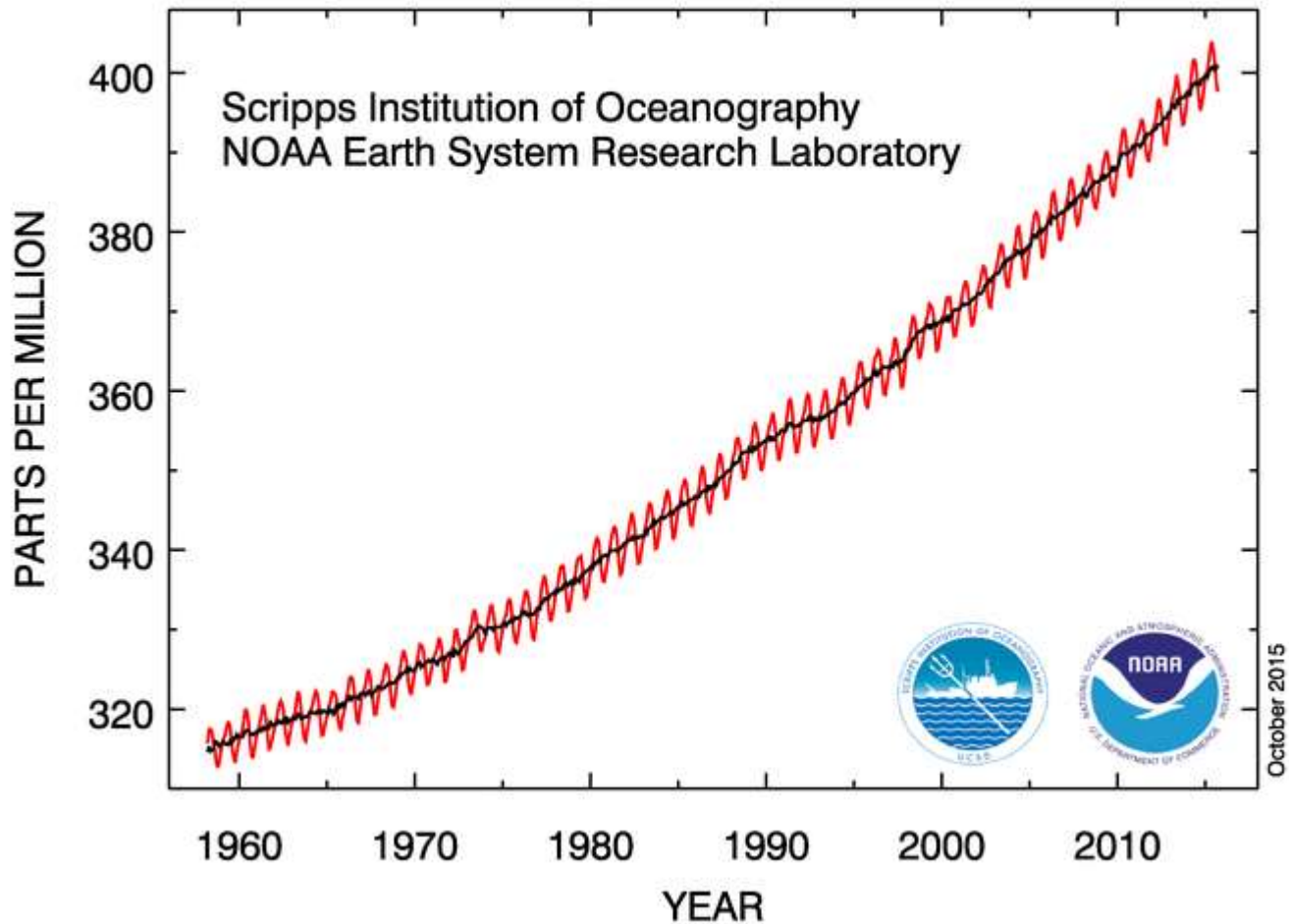
~55
ppmv



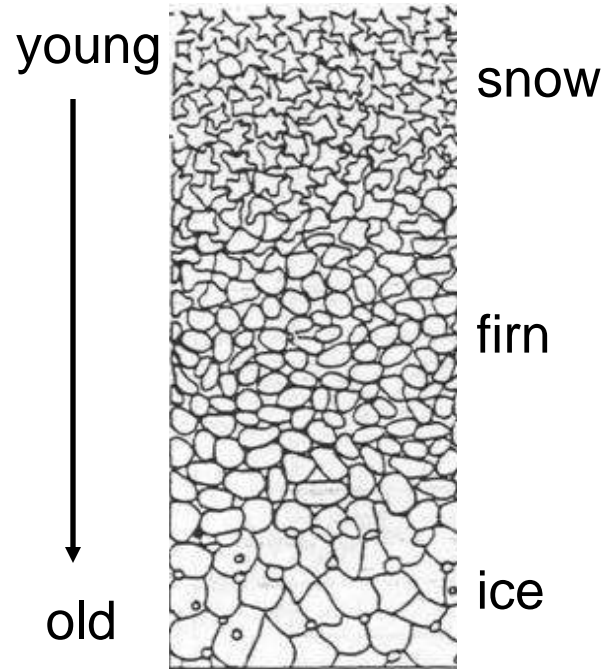
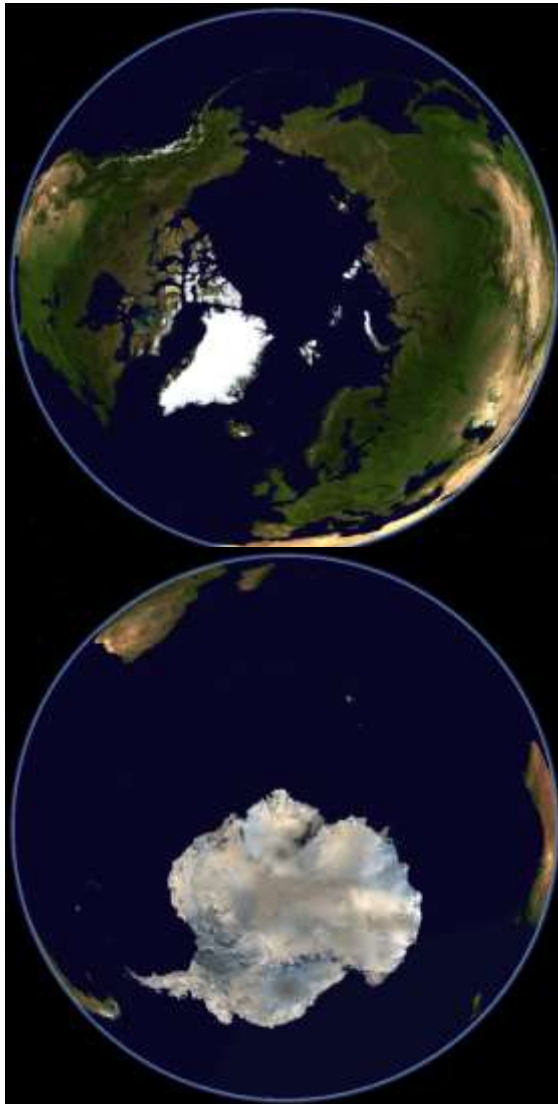
~90
ppmv

(Macke, 2004)

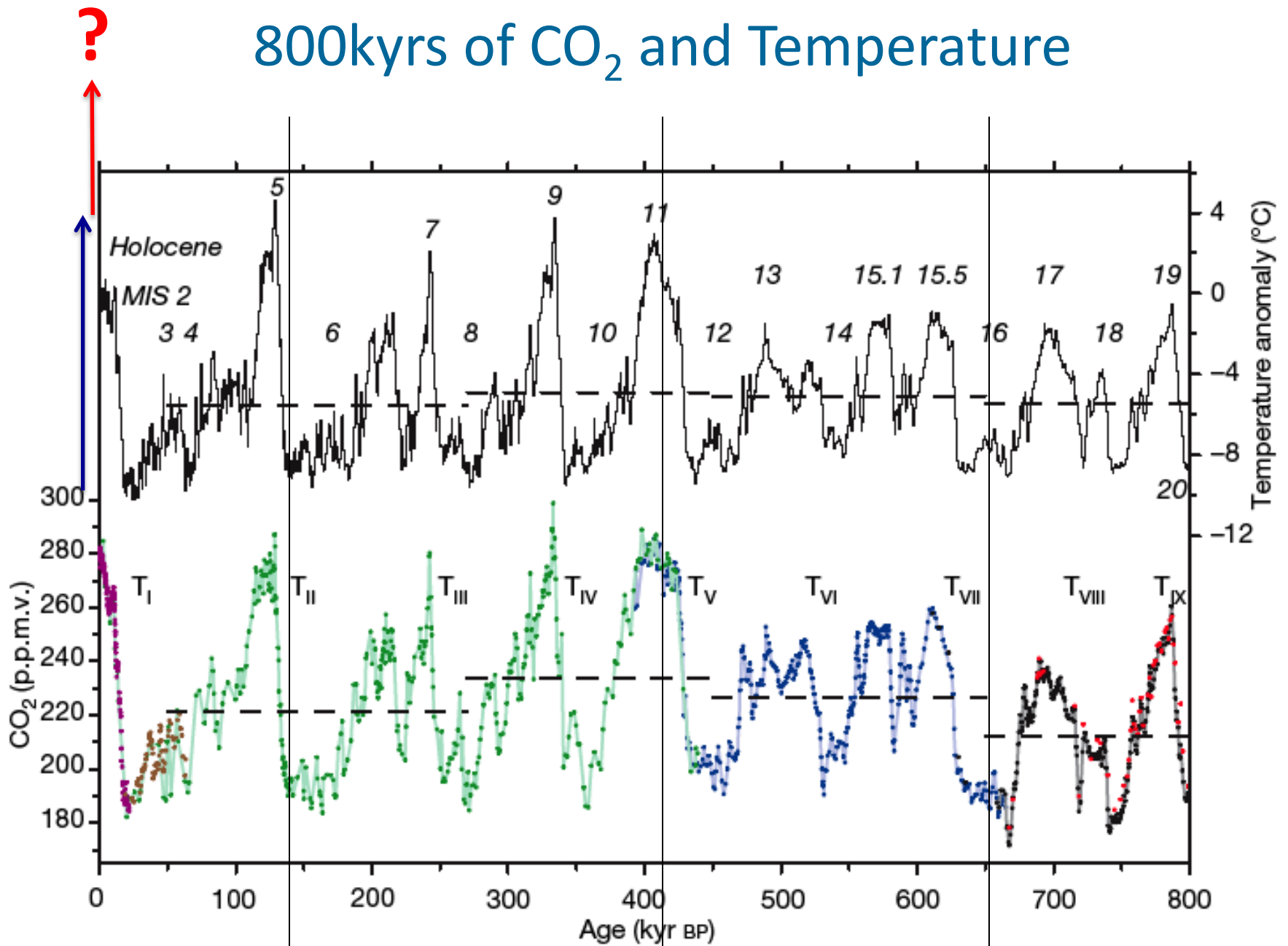
Atmospheric CO₂ at Mauna Loa Observatory



Ice cores as climate archives



800kyrs of CO₂ and Temperature



Barnola et al., 1987

Petit et al., 1999 Siegenthaler et al., 2005 Lüthi et al., 2008

