

Addendum A: Finding Courses

Accessing the correct information

FOR IPSU (International Programme Stellenbosch University) COURSES:

<http://www0.sun.ac.za/international/prospective-students/non-degree-seeking--short-term-students-1/i-want-to-enrol-at-su-1/exchange-programmes/course-information-1.html>

Please note that the IPSU booklet may be updated during the year but IPSU courses are usually repeated every semester. If you are unsure about the availability of an IPSU course, do not hesitate to send an email to interchange@sun.ac.za.

FOR MAINSTREAM COURSES:

Go to <http://www.sun.ac.za/english/faculty/> and click on the link for Calendar the relevant faculty – if this link does not work, do not hesitate to send an email to interchange@sun.ac.za

Understanding the Calendar

When making a choice for a mainstream course or module at the University of Stellenbosch, it is important to know what the different markers indicate to make it easier for you to make the right choice. This document gives a crash-course in deciphering the module codes so that you know exactly what you are choosing.

First things first: Language of instruction

Stellenbosch University Policy – Mainstream courses:

The following indicates the medium of instruction

(A) – Only Afrikaans

(E) – Only English

(T) – T option: 50/50

(A/E) – parallel: two class streams

* Take note that all academic material will be made available in English

INFO TO EXPLAIN COURSE DETAILS:

Each of the following abbreviations is used throughout the University to denote the stated combination of (i) type of instruction and (ii) length of time:

L lecture of 50 minutes;

P practical period of 50 minutes [for instance: 2P and 10P stand for totals of two practical periods and ten practical periods, respectively]; 3P a practical of maximally 170 minutes [this maximum is made up of 3 x 50 minutes, plus the two intervals of 10 minutes each]; S seminar of 50 minutes; T tutorial of 50 minutes; 2T tutorial of maximally 110 minutes [this maximum is made up of 2 x 50 minutes, plus the interval of 10 minutes].

Please familiarise yourself with the coding of mainstream and IPSU courses in order to make a course selection. The codes will tell you when the course is offered (in the case of

undergraduate course). The code consists of three digits. The first indicates the year-level of the course, the second indicates which semester the course is offered in (1, 2 and 3 indicates a course taught in the first semester; 4, 5 and 6 indicates a second semester course. 7 or 8 indicates a course that runs over a whole year). The third digit is irrelevant for your purposes. For example: English 348 is a third year course (3); offered in the second semester (4). The codes for postgraduate courses are more complicated, as the digits do not correspond to the semester it is offered in. For postgraduate courses, it is the easiest to go directly to the department's website to find out which courses are on offer and when they will be on offer and what the language of instruction will be.

Please see the example on the next page to help you.

To better explain the mentioned information, look at this extract from the yearbook from the Faculty of Science as an example

144	16	Fundamental Principles of Chemistry II	3L, 3P	A & E
Chemical equilibrium (both quantitative and qualitative) with applications in acid-base and precipitation reactions of aqueous solutions; an introductory study of organic compounds with a variety of functional groups; reaction mechanisms; stereochemistry; polymerisation. <i>C Chemistry 124</i>				
176	32	Introduction to Chemistry	3L, 3P	A & E
For students in the BSc (Extended Degree Programme). This module deals with the following themes: Classification of matter; atoms, molecules and ions; stoichiometry; reactions in aqueous solutions; atomic structure; chemical bonding; acid and bases; the periodic table. Examples that illustrate the importance and relevance of science as an everyday phenomenon.				
214	16	Organic Chemistry	3L, 3P	T
Reaction mechanisms, including nucleophilic addition and substitution, elimination, electrophilic addition, electrophilic aromatic substitution; stereochemistry. <i>PP Chemistry 114, 154</i>				
234	16	Inorganic Chemistry	3L, 3P	E
Periodic trends; structure and bonding in molecules; structure and bonding in solids; chemistry in solution; main group elements. Coordination chemistry: Introduction, types of ligands, nomenclature; isomerism in coordination compounds; different geometries; formation constants; crystal field theory. <i>PP Chemistry 114</i>				
254	16	Physical Chemistry	3L, 3P	A
Chemical thermodynamics; colligative properties; phase diagrams; reaction kinetics; electrochemistry. <i>P Mathematics 114, 144</i>				
264	16	Analytical Chemistry	3L, 3P	T
Introduction to chemical analysis; basic classical analytical chemistry; errors and uncertainty in analytical data; basic statistical methods; volumetric methods (acid-base, redox and complexometric analysis); solvent extraction; introduction to chromatographic separation; introduction to analytical molecular spectroscopy: fundamental principles and quantitative aspects of UV/visible spectrophotometry. <i>PP Chemistry 114, 154 and P Mathematics 114, 144 or P Mathematics (Bio) 124 or P Engineering Mathematics 115, 145</i>				
314	16	Analytical Chemistry	3L, 3P	E
Introduction to instrumental analysis. Error theory in quantitative chemical analysis, calibration in instrumental analysis and figures of merit. Introduction to atomic spectroscopy: atomic absorption and atomic emission spectroscopy for quantitative elemental analysis. Molecular spectroscopy: basic principles and application of ¹ H and ¹³ C nuclear magnetic resonance spectroscopy (NMR); introduction to infrared spectroscopy; introduction to analytical mass spectrometry; instrumental chromatographic methods. <i>P Chemistry 224</i>				

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 (A/E) – parallel: two class streams
 * Take note that all academic material will be made available in English

Indicates type of instruction and length of time (indicated per week):
 L lecture of 50 minutes;
 P practical period of 50 minutes [for instance: '2P' and '10P stand for totals of two practical periods and ten practical periods, respectively]; 3P a practical of maximally 170 minutes [this maximum is made up of 3 x 50 minutes, plus the two intervals of 10 minutes each]; S seminar of 50 minutes; T tutorial of 50 minutes; 2T tutorial of maximally 110 minutes [this maximum is made up of 2 x 50 minutes, plus the interval of 10 minutes].

This indicates the name of the module

This indicates the amount of credits for the course: 2 South African credits = 1 ECTS
 4 South African credits = 1 USA credit

The code consists of three digits. The first indicates the year-level of the course, the second indicates which semester the course is offered in (1, 2 and 3 indicates a course taught in the first semester; 4, 5 and 6 indicates a second semester course. 7 or 8 indicates a course that runs over a whole year). The third digit is irrelevant for your purposes. For example: English 348 is a third year course (3); offered in the second semester (4). The codes for postgraduate courses are more complicated, as the digits do not correspond to the semester it is offered in.

Fundamentals Principles of Chemistry II is thus a first year subject offered in the second semester, worth 16 credits and is taught both in Afrikaans and English. Instruction is given in 3 lectures and a practical of approximately 170 minutes per week.